

ORBITAL INTERACTIONS IN CHEMISTRY

Christen Atcheson

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Orbital Interaction Theory

Orbital Interactions in Chemistry 2nd Edition. Now in its second edition, this book explores organic, organometallic, inorganic, solid state, and materials chemistry, demonstrating how common molecular orbital situations arise throughout the whole chemical spectrum. MYUNG-HWAN WHANGBO.

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Balanced crystal orbital overlap population—a tool for analysing chemical bonds in solids. National Center for Biotechnology Information. As with H_2 the lowest energy atomic orbitals are the $1s'$ and $1s''$, and do not transform according to the symmetries of the molecule, while the symmetry adapted atomic orbitals .

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The symmetric combination—the bonding orbital—is lower in energy than the basis orbitals, and the antisymmetric combination—the antibonding Orbital Interactions in Chemistry higher. Necessarily this is an approximation, and highly accurate descriptions of the molecular electronic wave function do not have orbitals see configuration interaction. This means that rotation of the MO about the internuclear axis does not result in a phase change.

This means that rotation of the MO about the internuclear axis does not result in configuration interaction—general concepts: